**T04D05 - Formal Charge**

Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Formal charge is a way to “select” the more important resonance structure.

The formula for formal charge is:

**Formal charge = Number of valence electrons – [ ½ number of shared electrons + number of unshared valence electrons]**

a. Sulfur trioxide

b. Laughing gas, N2O

c. Cyanate ion, NCO-1

d. Carbon dioxide

e. Carbon disulfide

f. Thiocyanate ion, SCN-1